

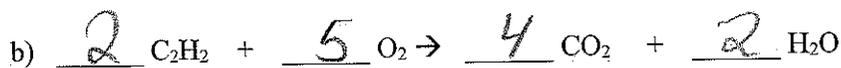
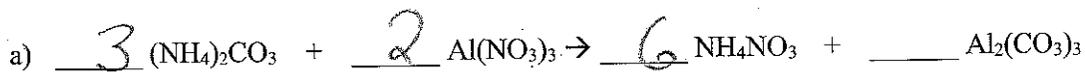
Key

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|----------------------------|--|----------------------------------|---|
| 1. Copper (II) bromide | <u>CuBr₂</u> | 14. Potassium permanganate | <u>KMnO₄</u> |
| 2. Aluminum sulfate | <u>Al₂(SO₄)₃</u> | 15. Mercury (I) oxide | <u>Hg₂O</u> |
| 3. Potassium nitrate | <u>KNO₃</u> | 16. Lead (II) chloride | <u>PbCl₂</u> |
| 4. Iron (II) carbonate | <u>FeCO₃</u> | 17. Hydroiodic acid | <u>HI (aq)</u> |
| 5. Lead (II) phosphate | <u>Pb₃(PO₄)₂</u> | 18. Iron (II) sulfite | <u>Fe(SO₃)</u> |
| 6. Diphosphorous pentoxide | <u>P₂O₅</u> | 19. Magnesium nitrate | <u>Mg(NO₃)₂</u> |
| 7. Nickel (II) nitrate | <u>Ni(NO₃)₂</u> | 20. Sodium phosphate | <u>Na₃PO₄</u> |
| 8. Ammonium dichromate | <u>(NH₄)₂Cr₂O₇</u> | 21. Magnesium sulfite | <u>MgSO₃</u> |
| 9. Zinc sulfate | <u>ZnSO₄</u> | 22. Calcium sulfate dihydrate | <u>CaSO₄ · 2H₂O</u> |
| 10. Tin (II) chloride | <u>SnCl₂</u> | 23. Calcium chloride dihydrate | <u>CaCl₂ · 2H₂O</u> |
| 11. Ammonium carbonate | <u>(NH₄)₂CO₃</u> | 24. Barium chloride tetrahydrate | <u>BaCl₂ · 4H₂O</u> |
| 12. Nickel (II) acetate | <u>Ni(C₂H₃O₂)₂</u> | 25. Sulfurous acid | <u>H₂SO₃ (aq)</u> |
| 13. Sodium chromate | <u>Na₂CrO₄</u> | | |

B. Provide the name for each of the following:

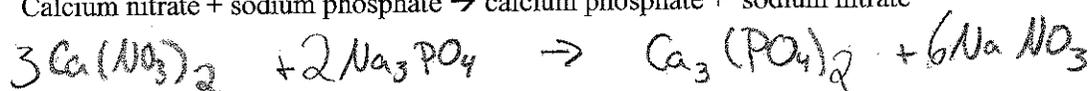
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|---|--------------------------------|---|---------------------------------------|
| 1. HgF ₂ | <u>Mercury (II) fluoride</u> | 17. Fe ₂ (SO ₄) ₃ | <u>Iron (III) Sulfate</u> |
| 2. KCl | <u>Potassium chloride</u> | 18. Mg(CN) ₂ | <u>Magnesium cyanide</u> |
| 3. KMnO ₄ | <u>Potassium Permanganate</u> | 19. HNO ₃ | <u>Nitric acid</u> |
| 4. KClO ₃ | <u>Potassium Chlorate</u> | 20. Na ₂ S · 5H ₂ O | <u>Sodium Sulfide · Pentahydrate</u> |
| 5. NH ₄ MnO ₄ | <u>Ammonium Permanganate</u> | 21. Fe(NO ₂) ₂ | <u>Iron (II) Nitrite</u> |
| 6. Ba ₃ (PO ₄) ₂ | <u>Barium Phosphate</u> | 22. HCl | <u>hydrochloric acid</u> |
| 7. Fe ₂ O ₃ | <u>Iron (III) oxide</u> | 23. SbCl ₃ | <u>Antimony (III) chloride</u> |
| 8. CoF ₃ | <u>Cobalt (III) Fluoride</u> | 24. P ₄ O ₁₀ | <u>tetra phosphorus decaoxide</u> |
| 9. Ca(ClO ₃) ₂ | <u>Calcium chlorate</u> | 25. Ba(OH) ₂ | <u>Barium hydroxide</u> |
| 10. HC ₂ H ₃ O ₂ | <u>acetic acid</u> | 26. HF | <u>hydrofluoric acid</u> |
| 11. LiI | <u>Lithium iodide</u> | 27. Fe(NO ₂) ₃ · 6H ₂ O | <u>Iron (III) Nitrite hexahydrate</u> |
| 12. Al ₂ (SO ₄) ₃ | <u>Aluminum Sulphate</u> | | |
| 13. HBr | <u>Hydro Bromic acid</u> | | |
| 14. CrCl ₃ | <u>Chromium (III) chloride</u> | | |
| 15. H ₃ PO ₄ | <u>Phosphoric acid</u> | | |
| 16. LiMnO ₄ | <u>Lithium Permanganate</u> | | |

1. For each of the following fully **balance** with lowest ratios:

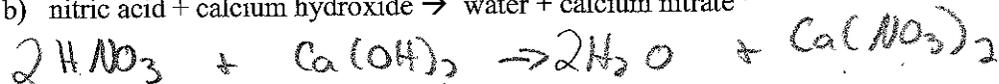


2. For each of the following equations, create a **balanced** equation:

a) Calcium nitrate + sodium phosphate \rightarrow calcium phosphate + sodium nitrate



b) nitric acid + calcium hydroxide \rightarrow water + calcium nitrate



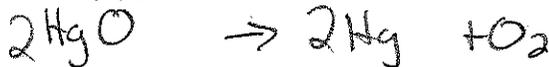
c) potassium iodide + lead (II) nitrate \rightarrow potassium nitrate + lead (II) iodide



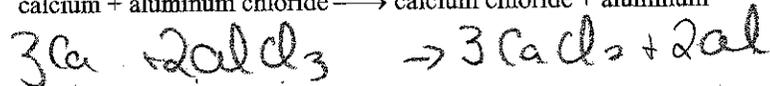
d) iron (III) oxide + carbon \rightarrow carbon monoxide (g) + iron



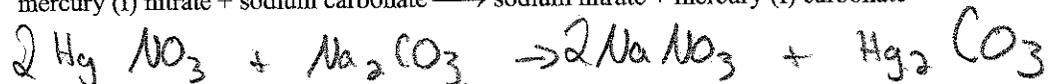
e) mercury (II) oxide \rightarrow mercury + oxygen (g)



f) calcium + aluminum chloride \rightarrow calcium chloride + aluminum



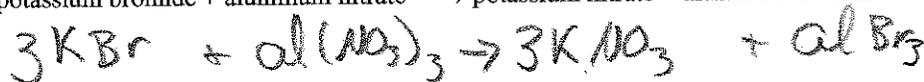
g) mercury (I) nitrate + sodium carbonate \rightarrow sodium nitrate + mercury (I) carbonate



h) rubidium + acetic acid (CH_3COOH) \rightarrow rubidium acetate + hydrogen (g)



i) potassium bromide + aluminum nitrate \rightarrow potassium nitrate + aluminum bromide



j) calcium phosphate + aluminum sulfate \rightarrow calcium sulfate + aluminum phosphate

